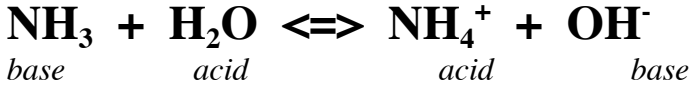
Brønsted Acids and Bases

A **Brønsted Acid** will donate a proton, while a **Brønsted Base** will accept a proton



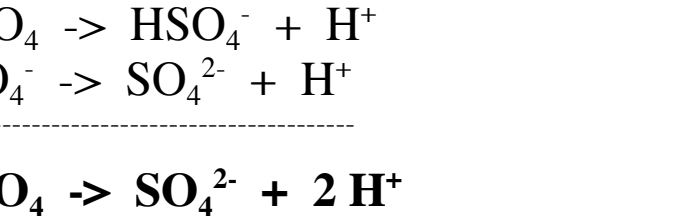
A base produces a **conjugate acid** upon acceptance of a proton

An acid produces a **conjugate base** upon release of a proton

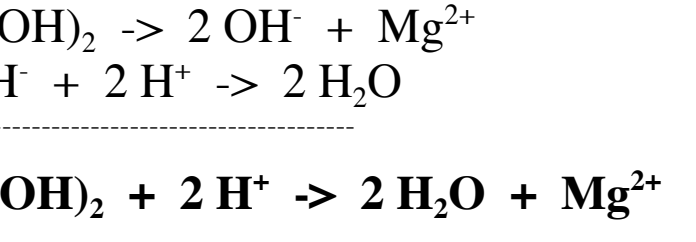
Acids do not create conjugate acids, and bases do not create conjugate bases

Some acids and bases are **polyprotic** - more than one proton can be released or accepted.

Example: Sulfuric acid is a diprotic acid



Example: Magnesium hydroxide is a diprotic base



Strong Acids and **Strong Bases** are completely ionized.
Weak Acids and **Weak Bases** have equilibrium constants, **K**, less than one. Use your textbook as a source of K values for acids and bases.