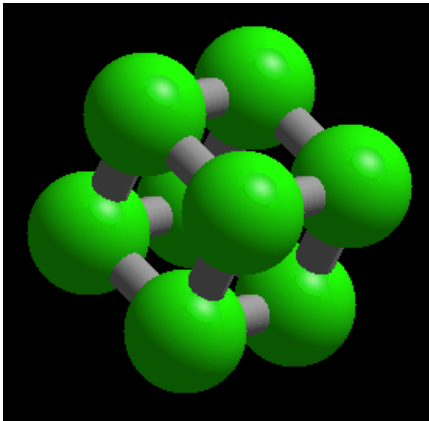
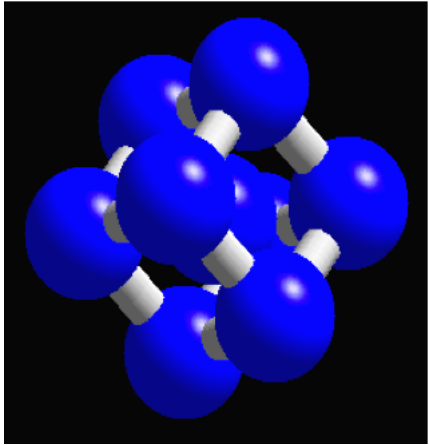



## Cubic Unit Cells Guide for Metallic Solids

Type	Atoms per Unit Cell	Edge - Radius Equation	Picture
Simple Cubic (SC)	1	Edge = 2*radius	
Body Centered Cubic (BCC)	2	Edge = $\frac{4 * \text{radius}}{\sqrt{3}}$	
Face Centered Cubic (FCC)	4	Edge = $\frac{4 * \text{radius}}{\sqrt{2}}$	

Volume = (edge)<sup>3</sup> for simple cubic unit cells  
 1 pm = 10<sup>-12</sup> m, 1 Å = 10<sup>-10</sup> m