

Building Lewis Dot Structures

1) **Decide on the central atom**

Central atom has lowest electron affinity

Never Hydrogen! H only terminal atom

2) **Count Valence Electrons & Electron Pairs**

Valence Electrons = Group Number

Electron Pairs = (Valence Electrons) / 2

3) **Form sigma bond(s) between central atom and surrounding atom(s)**

4) **Remaining electrons form lone pairs**

Use the *Octet Rule* to assign electrons

5) **If an atom(s) does not have octet, create pi bond(s) using lone pairs on adjacent atoms**