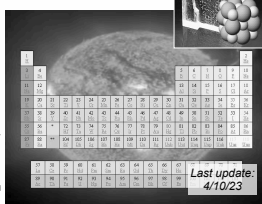
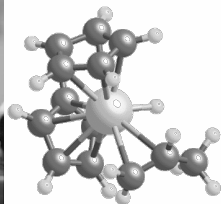




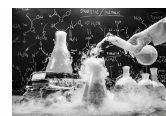
Welcome to Chemistry 221!

<http://mhchem.org/221>

Get the CH 221 Companion before lab!



CH 221: Lectures and Labs



Lectures: MWF from 9 - 9:50 AM in AC 1303 (this room)

- Lectures recorded, available soon afterwards
- Lecture notes to print available (under "Problem Sets and Handouts", mhchem.org/221) and in **Chemistry 221 Companion** (get it!)

Labs (Section 01): Mondays from 1:10 - 5 PM

- Start in room AC 1303
- Move to AC 2507 ("the lab") around 3 PM
- For first day, bring a printed copy of the "Eight Bottles" Lab (mhchem.org/221) and your calculator
- Some labs will require safety glasses (Dollar store ok)

MAR

...more on Monday afternoon

The Nature of Chemistry



What is Chemistry?

- "Keme" (earth)
- "Kehmeia" (transmutation)
- "Al-Khemia" (Arabic)
- "alchemy" (Europe's Dark Age)
- "chymistry" (Boyle's 1661 publication)
- M. • "chemistry" (modern)

Khemeia - ancient Egyptian processes for embalming the dead, later extended to metallurgy

Khemeia (and later chemistry) seen as "occult" by laymen, extended to modern age

What is Matter?

The Nature of Chemistry



How does Matter Change?

How does Matter Interact?



Why Study Chemistry?



MAR

The Art (?) of Chemistry



Chemistry and Art!?

Dr. Roald Hoffman,
1981 Nobel Prize in Chemistry

Stick to the chemistry,
Roald!

"There was no question that the reaction worked but transient colors were seen in the slurry of sodium methoxide in dichloromethane and we got a whole lot of products for which we can't sort out the kinetics the next slide show will show the most important part very rapidly within two minutes and I forgot to say on further warming we get in fact the ketone..."

- Organic - carbon, nitrogen, oxygen
- Inorganic - metals, everything "non-carbon"
- Analytical - Spectroscopy, "how much", "what kind"
- Physical - measurement, where physics meets chemistry
- Biochemical - the chemistry of life
- many others!

The Branches of Chemistry

MAR

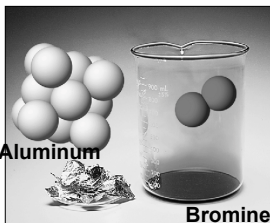
MAR

The Language of Chemistry

CHEMICAL ELEMENTS - pure substances that cannot be decomposed by ordinary means to other substances.



Sodium



Aluminum

Bromine

MAR

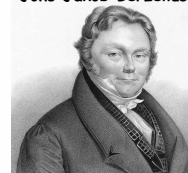
The Language of Chemistry

The elements, their names, and symbols are given on the **PERIODIC TABLE**

Berzelius - first to use letter symbols for atoms

How many elements are there?

Jöns Jakob Berzelius

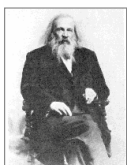


pre-Biblical elements: Au (*sun*), Ag (*moon*), Cu (*Venus*), Fe (*Mars*), Sn (*Jupiter*), Pb (*Saturn*), Hg (*Mercury*), S, C

MAR

The Periodic Table

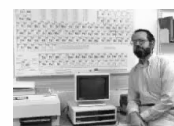
Periodic table originally organized by mass, now by atomic number



Dmitri Mendeleev (1834 - 1907)
Predicted Ga, Ge, Sc and Tc!

Dr. Frank DiSalvo (Cornell University)

"On the importance of the periodic table"

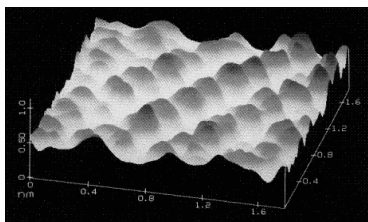
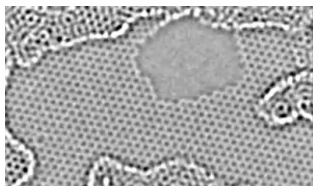


Number of compounds possible is virtually limitless!!!

MAR

An **atom** is the smallest particle of an element that has the chemical properties of the element.

Real time carbon atoms from TEAM 0.5 / NCEM



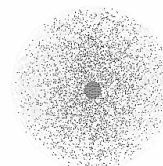
Copper atoms on a silica surface
Distance across = 1.8 nanometer (1.8×10^{-9} m)

MAR

MAR

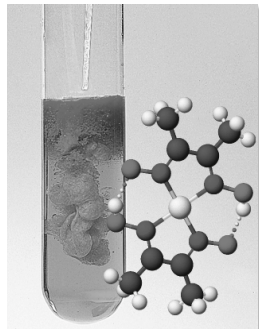
The Atom

An atom consists of a nucleus (of protons and neutrons) and electrons in space about the nucleus.



— Electron cloud
— Nucleus

CHEMICAL COMPOUNDS are composed of atoms and so can be decomposed to those atoms.

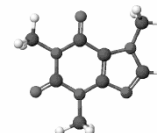
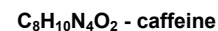
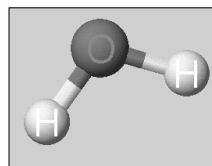


The red compound is composed of
 nickel (Ni) (silver)
 carbon (C) (black)
 hydrogen (H) (white)
 oxygen (O) (red)
 nitrogen (N) (blue)

This type of compound is an ionic compound - unshared electrons

A MOLECULE is the smallest unit of a compound that retains the chemical characteristics of the compound.

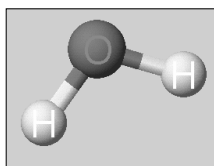
Composition of molecules is given by a **MOLECULAR FORMULA**



MAR

A MOLECULE is the smallest unit of a compound that retains the chemical characteristics of the compound.

Composition of molecules is given by a **MOLECULAR FORMULA**



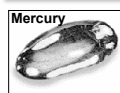
Water and caffeine are examples of covalent compound - shared electrons

MAR

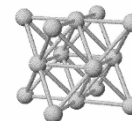
The Nature of Matter



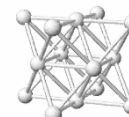
Gold



Mercury



Gold



Mercury

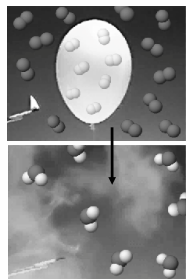
Chemists are interested in the nature of matter and how this is related to its atoms and molecules.

We can explore the **MACROSCOPIC** world - what we can see - to understand the **ATOMIC** world - what we cannot see - using **SYMBOLS**.

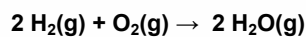
MAR

A Chemist's View

Macroscopic



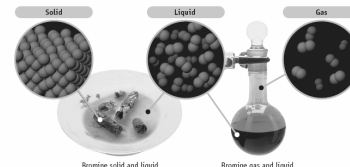
Atomic



Symbolic

MAR

STATES OF MATTER



SOLIDS: rigid shape, fixed volume, reasonably well understood.

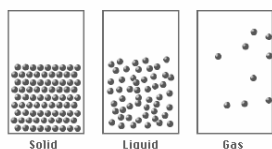
LIQUIDS: no fixed shape, may not fill a container completely, not well understood.

GASES: expand to fill their container, good theoretical understanding.

also **PLASMA** - more in CH 222!

KINETIC NATURE OF MATTER

Matter consists of atoms and molecules in motion.



Kinetic Molecular Theory describes solids, liquids and gases

Test Monkeys? Er, sorry, Student volunteers?!?

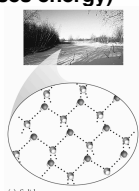
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STATES OF MATTER & ENERGY

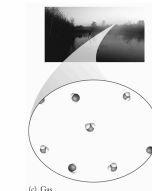
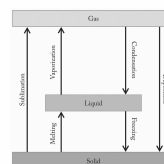
Energy of transitions important to scientists:

solid -> liquid -> gas: endothermic (takes energy)

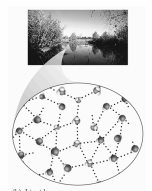
gas -> liquid -> solid: exothermic (releases energy)



MAR



(c) Gas

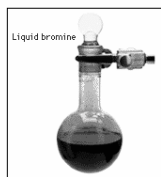


(b) Liquid

Physical Properties

Physical properties can be observed and measured without changing the composition of a substance. They include:

- color
- melting and boiling point
- odor



Some physical changes would be

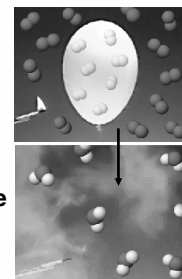
- boiling of a liquid
- melting of a solid
- dissolving a solid in a liquid to give a homogeneous mixture - a SOLUTION.

MAR

Chemical Properties and Chemical Change

Burning hydrogen (H₂) in oxygen (O₂) gives H₂O.

Chemical change or chemical reaction involves the transformation of one or more atoms or molecules into one or more different molecules.



MAR

Chemical Properties and Physical Properties

Physical properties do not change the composition of the substance

Chemical properties change the composition of the substance



MAR

Physical Properties

Physical properties useful in separating compounds and elements

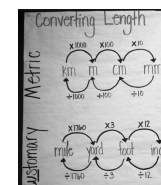
- density
- melting and boiling point
- magnetism

Physical and chemical properties require units - need

METRIC SYSTEM!

See the Metric Guide

MAR



The Metric System



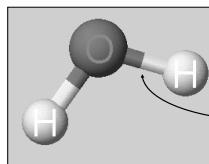
Use handouts, examples in book to learn the metric system

MAR

Units of Length / Conversions

- 1 kilometer (km) = 10^3 meters (m)
- 1 centimeter (cm) = 10^{-2} meters (m)
- 1 millimeter (mm) = 10^{-3} meters (m)
- 1 micrometers (μm) = 10^{-6} meters (m)
- 1 nanometer (nm) = 10^{-9} meters (m)

Know these five metric conversions!



O-H distance =
 9.4×10^{-11} m
 9.4×10^{-9} cm
 9.4×10^{-5} μm
 0.094 nm

MAR

Density: the ratio of a substance's mass (grams) to its volume (mL, cm³)



Substances layer themselves according to their density: superposition

Density used to separate materials



Gold
 $d = 19.3 \text{ g/cm}^3$



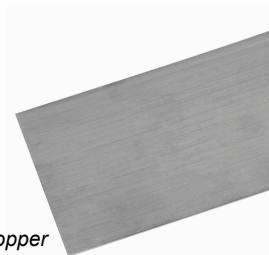
Mercury
 $d = 13.6 \text{ g/cm}^3$

MAR

Density Problem

Problem: A piece of copper has a mass of 57.54 g. It is 9.36 cm long, 7.23 cm wide, and 0.95 mm thick. Calculate density (g/cm³).

$$\text{Density} = \frac{\text{mass (g)}}{\text{volume (cm}^3\text{)}}$$



Sheet of copper

MAR

Density Problem

Problem: A piece of copper has a mass of 57.54 g. It is 9.36 cm long, 7.23 cm wide, and 0.95 mm thick. Calculate density (g/cm³).

SOLUTION

1. Get dimensions in common units.

$$0.95 \text{ mm} \cdot \frac{1 \text{ cm}}{10 \text{ mm}} = 0.095 \text{ cm}$$

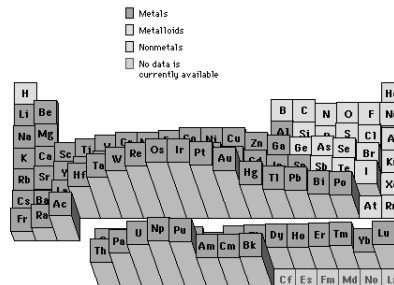
2. Calculate volume in cubic centimeters.

$$(9.36 \text{ cm})(7.23 \text{ cm})(0.095 \text{ cm}) = 6.4 \text{ cm}^3$$

6.42891...

3. Calculate the density.

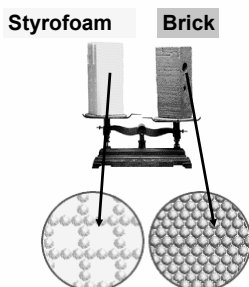
Relative Densities of the Elements



MAR

DENSITY

Notice that density is an **INTENSIVE** property of matter
INTENSIVE - does not depend on quantity of matter (density, boiling point, etc.)
 Contrast with **EXTENSIVE** - depends on quantity of matter. Examples include mass and volume



MAR

PROBLEM: Mercury (Hg) has a density of 13.6 g/cm³. What is the mass of 95 mL of Hg? (454 g = 1 lb)



Solve the problem using **DIMENSIONAL ANALYSIS** - see the *Dimensional Analysis* and *Factor Label* handouts on the Web

MAR

PROBLEM: Mercury (Hg) has a density of 13.6 g/cm³. What is the mass of 95 mL of Hg? (454 g = 1 lb)

First, note that **1 cm³ = 1 mL**

Then, use dimensional analysis to calculate mass.

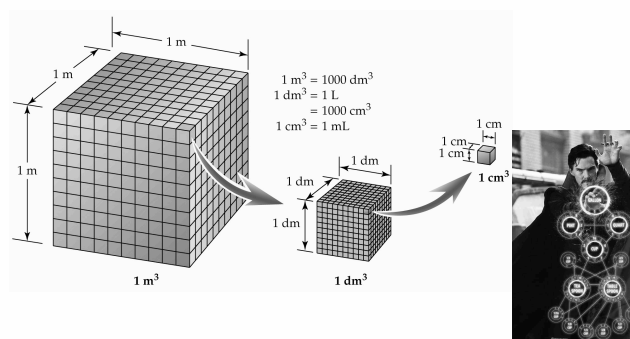
$$95 \text{ cm}^3 \cdot \frac{13.6 \text{ g}}{1 \text{ cm}^3} = 1.3 \times 10^3 \text{ g}$$

What is the mass in pounds?

$$1.3 \times 10^3 \text{ g} \cdot \frac{1 \text{ pound}}{454 \text{ g}} = 2.9 \text{ lb}$$

MAR

Volume, Liters and Cubed Lengths



MAR

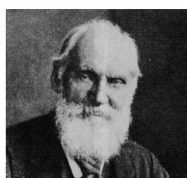
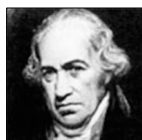
Temperature Scales

- Fahrenheit (°F)
- Celsius (°C)
- Kelvin (K)



Anders Celsius
1701-1744

Daniel Fahrenheit
1686-1736



Lord Kelvin
(William Thomson)
1824-1907

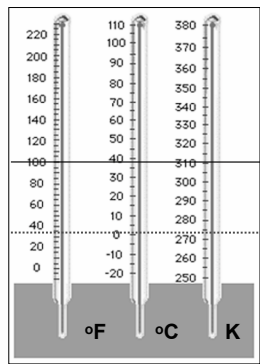
MAR

MAR

Temperature Scales

	Fahrenheit	Celsius	Kelvin
Boiling Point of water	212	100	373.15
	↑ 180°	↑ 100°	↑ 100°
Freezing Point of water	32	0	273.15

Notice that 1 Kelvin degree = 1 degree Celsius
 Difference between Celsius temperatures and Kelvin temperatures the same!



Temperature Scales

$$100\text{ }^{\circ}\text{F} = 38\text{ }^{\circ}\text{C} = 311\text{ K}$$

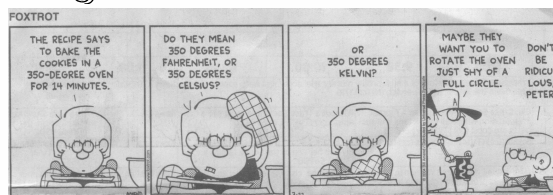
$$32\text{ }^{\circ}\text{F} = 0\text{ }^{\circ}\text{C} = 273\text{ K}$$

Note:
no negative K!
no degree K symbol!

MAR



Temperature Scales



MAR



Calculations Using Temperature

Chemistry calculations *generally* require temperatures in Kelvins (K)

$$T\text{ (K)} = T\text{ (}^{\circ}\text{C)} + 273.15$$

$$\text{Body temp} = 37\text{ }^{\circ}\text{C} + 273 = 310.\text{ K}$$

$$\text{Liquid nitrogen} = 77\text{ K} - 273 = -196\text{ }^{\circ}\text{C}$$

MAR



Calculations Using Temperature

Occasionally need Fahrenheit (F) values

Convert using Celsius scale

$$T\text{ (}^{\circ}\text{F)} = \frac{9}{5} T\text{ (}^{\circ}\text{C)} + 32.00$$

$$\text{Liquid He} = 4.2\text{ K} - 273.15 = -269.0\text{ }^{\circ}\text{C}$$

$$T\text{ (}^{\circ}\text{F)} = \frac{9}{5} (-269.0\text{ }^{\circ}\text{C)} + 32.00 = -452.1\text{ }^{\circ}\text{F}$$



MAR

UNITS OF MEASUREMENT

We make **QUALITATIVE** observations of reactions - changes in color and physical state.

We also make **QUANTITATIVE MEASUREMENTS**, which involve numbers and amount.

Use SI units - based on the metric system

length	(meter, m)
mass	(kilogram, kg)
time	(second, s)

MAR



Don't stand next to the dart board, especially with poor precision!
MAR

Accuracy and Precision

Measurements affected by accuracy and precision.



Accurate and Precise

Accuracy versus Precision

Accuracy refers to the proximity of a measurement to the true value of a quantity

Accuracy determined by % error

Precision refers to the proximity (reproducibility) of several measurements to each other.

Determined by average deviation or parts per thousand



MAR

Experimental Error

Average deviation:

Step 1: find the absolute value of the difference between each measurement and the average.

Step 2: find the summation of all the deviations and divide by the total number of measurements.

Standard deviation (not used in CH221):

$$\text{Standard deviation} = \sqrt{\frac{\text{sum of squares of deviations}}{(\# \text{ of deviations} - 1)}}$$

ppt (parts per thousand):

$$\text{ppt} = \frac{\text{average deviation}}{\text{average}} \times 1000$$

Percent error:

$$\% \text{ error} = \frac{\text{experimental value} - \text{accepted value}}{\text{accepted value}} \times 100$$

MAR

Experimental Error - Example

Trial #	Boiling Point (°C)	Average (°C)	Deviations (°C)	Ave. Dev. (°C)
1	11.23	11.19	0.04	0.06
2	11.09		0.10	
3	11.27		0.08	
4	11.16		0.03	

Average Deviation = 0.06 °C (11.19 ± 0.06 °C)

ppt = (0.06 °C / 11.19 °C) x 1000 = 5 ppt

If the literature (accepted) value was 11.25 °C,

% error = (11.19 °C - 11.25 °C / 11.25 °C) x 100 = -0.5%

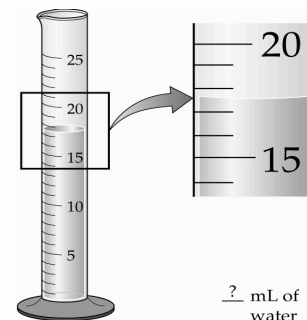
sometimes %error is absolute value (always positive)

MAR

Measurement and Significant Figures

Every experimental measurement, *no matter how precise*, has a degree of uncertainty because there is a limit to the number of digits that can be determined.

Need mathematical system - **SIGNIFICANT FIGURES - very important**, see Chapter One in text and Handout



MAR

Measurement and Significant Figures

To indicate the precision, recorded values should use all the digits known with certainty *plus* one additional estimated digit

Estimated ("doubtful") digit usually considered *uncertain* by plus or minus 1 (± 1)

The total number of digits used to express such a measurement is called the *number of significant figures (sig figs)*.

Ex: 65.07 g - four sig figs, 7 "doubtful"

Ex: 54.70318 g - seven sig figs, 8 "doubtful"

MAR

Rules for Determining Significant Figures

1. Zeroes in the middle of a number are significant. 69.08 has four sig figs.
2. Zeroes at the beginning of a number are not significant. 0.0089 has two sig figs (8 and 9).
3. Zeroes at the end of a number and after the decimal point are significant. 2.50 has three sig figs. 25.00 has four sig figs.
4. Zeroes at the end of a number and before the decimal point will be significant *only* with a decimal placeholder (period). 1500 has two sig figs, but 1500. has four sig figs.
5. Exact conversions (Definitions) have *infinite* sig figs (ex: 60 s/1 min, 10 mm/1 cm).
6. **STUDY! PRACTICE! IMPORTANT!**

MAR

Scientific Notation

Scientific Notation is a convenient way to write very small or large numbers

Know how your calculator displays scientific notation (and also "regular" notation!)

Written as a product of a number between 1 and 10, times the number 10 raised to a power. *Examples:*

$$215. = 2.15 \times 10^2$$

Decimal point is moved two places to the left, so exponent is 2.

$$1.56 \times 10^{-8} = 0.000\,000\,015\,6$$

Negative exponent of -8, so decimal point is moved to the left eight places.

Always use proper scientific notation when reporting answers in lab, quizzes, etc.

MAR

See: [Scientific Notation Handout](#) & [Scientific Notation Handout #2](#)

MAR

Rounding Numbers with Sig Figs

Calculators produce large numbers in calculations *although* the number of sig figs is good only to a few numbers, *much less* than the calculator has produced

In this case the large number must be rounded off to a smaller number keeping only significant figures.

Rules for Rounding off Numbers

Rule 1 (For multiplication and division): The answer cannot have more significant figures than either of the original numbers.

$$\frac{\begin{array}{l} \text{Three significant} \\ \text{figures} \end{array} \begin{array}{l} 278 \text{ mi} \\ 11.70 \text{ gal} \end{array}}{\begin{array}{l} \text{Four significant} \\ \text{figures} \end{array}} = \begin{array}{l} \text{Three significant} \\ \text{figures} \end{array} 23.8 \text{ mi/gal}$$

Actual value: 23.76068.....

MAR

MAR

Rules for Rounding off Numbers

Rule 2 (For addition and subtraction): The final number must stop at the *largest doubtful digit*.

Volume of water at start	→	3.18? ?? L	} 3.18: 8 is the "doubtful digit", it stops at the hundredths spot
Volume of water added	→	+ 0.01315 L	
Total volume of water	→	3.19? ?? L	

0.01315: 5 is the "doubtful digit", it stops at the hundredth thousandths spot

Final answer is 3.19 L:
Answer stops at largest "doubtful digit" (hundredths vs. hundredth thousandths)

Actual value: 3.19315

Rules for Rounding off Numbers

Rule 3: Once you decide how many numbers to keep, you *may* need to round off your answer:

If the first digit you remove is between 0 and 4, drop it and all remaining digits.

If the first digit you remove is between 5 and 9, round the number up by adding 1 to the digit to the left of the one you drop

Example: 2.4271 becomes 2.4 when rounded to two significant figures

Example: 4.5816 becomes 4.6 when rounded to two significant figures

Use these rules when rounding numbers - do not use Textbook's weird "even rounding with a 5"!

MAR

MAR

Mass Percentages in Chemistry

Often see "30% lead, 70% oxygen"

This means that in 100 grams of the substance

30 grams will be lead

70 grams will be oxygen

In one gram of the substance,

0.30 grams will be lead

0.70 grams will be oxygen



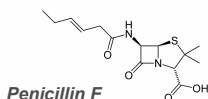
Mass Percentages in Chemistry

Example: Penicillin F is 53.829% carbon. How much carbon in 75 g of Penicillin F?

Solution

75 g Penicillin F* (53.829 g C / 100 g Penicillin F) = 40. g carbon (40.37175)

Note that volume percentages occasionally seen, but not often in our class

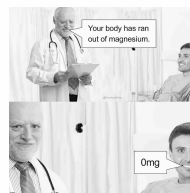


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End of Chapter One

See also:

- [Chapter One Study Guide](#)
- [Chapter One Concept Guide](#)
- [Math \("Chapter Guide Zero"\) Concept Guide](#)
- [Important Equations \(following this slide\)](#)
- [End of Chapter Problems \(following this slide\)](#)



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Important Equations, Constants, and Handouts from this Chapter:

metric prefixes:
nano (n) = 10⁻⁹
micro (μ) = 10⁻⁶
milli (m) = 10⁻³
centi (c) = 10⁻²
kilo (k) = 10³

$$\text{Density} = \frac{\text{mass (g)}}{\text{volume (cm}^3\text{)}}$$

$$T (\text{K}) = T (^\circ\text{C}) + 273.15$$

1 cm³ = 1 mL

significant figures!!!

mass percentages

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End of Chapter Problems: Test Yourself

1. $32.32 - 23.2 = \underline{\hspace{1cm}}$
2. $32.4 \cdot 37.31 = \underline{\hspace{1cm}}$
3. $4.311 / 0.07 = \underline{\hspace{1cm}}$
4. Convert 37.0 C to K.
5. Convert 253.6 mL to cm³.
6. Convert 24 m³ to cm³.
7. $235.05 + 19.6 + 2 = \underline{\hspace{1cm}}$
8. $58.925 - 19 = \underline{\hspace{1cm}}$
9. $2.19 \times 4.2 = \underline{\hspace{1cm}}$
10. $4.311 \div 0.07 = \underline{\hspace{1cm}}$
11. The platinum-containing cancer drug cisplatin contains 65.0% platinum. If you have 1.53 g of the compound, what mass of platinum (in grams) is contained in this sample?
12. The anesthetic procaine hydrochloride is often used to deaden pain during dental surgery. The compound is packaged as a 10.0% solution (by mass; $d = 1.0 \text{ g/mL}$) in water. If your dentist injects 0.50 mL of the solution, what mass of procaine hydrochloride (in milligrams) is injected?

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End of Chapter Problems: Answers

1. 9.1
2. 1210
3. 60
4. 310.2 K
5. 253.6 cm³
6. $2.4 \times 10^7 \text{ cm}^3$
7. 257
8. 40.
9. 9.2
10. 60
11. 0.995 g Pt
12. 50. mg

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