Dissolve, Dissociate and Electrolyte Guide

$$AgCl(s)$$
 $\xrightarrow{H_2O}$ No reaction, does not dissolve or dissociate silver(1) chloride

NaCl_(s)
$$\xrightarrow{H_2O}$$
 NaCl_(aq) \xrightarrow{time} Na⁺_(aq) + Cl⁻_(aq) sodium chloride $\xrightarrow{H_2O}$ dissolves

$$C_6H_{12}O_{6(s)} \xrightarrow{H_2O} C_6H_{12}O_{6(aq)} \xrightarrow{time} C_6H_{12}O_{6(aq)}$$
sugar

 $dissolves$

No dissociation nonelectrolyte does not conduct electricity

Partial dissociation
weak electrolyte
partial conductor of electricity