"Quick & Dirty" Rules for Ionic Compounds

- I. Generally, metals form ionic compounds
- II. Nonmetals *generally* form *ionic* compounds when allowed to react with metals
- III. Metalloids are difficult to predict
- IV. The farther apart the elements are in the periodic table, the *better the chances* the compound will be ionic

$$Ex: Na^+ + Cl^- \rightarrow NaCl$$

$$Ex: Cd^{2+} + Te^{2-} \rightarrow CdTe (?)$$