CH 151 Problem Set #3

Complete problem set on separate pieces of paper showing all work, circling final answers, etc.

Covering: **Chapter Two and Chapter Three**
Important Tables and/or Constants: 1 mol = 6.022×10^{23}

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1. Name each of the following fixed-charge binary ionic compounds.
    a. BeS
    b. $GaCl_3$
    c. CaO
    d. Cd_3P_2
2. Name each compound in the following pairs of variable-charge binary ionic compounds.
    a. $SnCl_4$ and $SnCl_2$
    b. FeS and Fe_2S_3
    c. Cu_3N and Cu_3N_2
    d. NiO and Ni_2O_3
3. Name each of the following binary ionic compounds.
    a. Au_2O
    b. Ag_2O
    c. CuCl
    d. KCl
4. Write chemical formulas for the following variable-charge binary ionic compounds, given their names.
    a. iron(II) chloride
    b. iron(III) chloride
    c. gold(I) oxide
    d. gold(I) iodide
5. Write chemical formulas for both ions in each of the following pairs of polyatomic ions.
    a. nitrate and nitrite
    b. chlorate and perchlorate
    c. cyanide and azide
    d. phosphate and phosphite
6. In which of the following pairs of compounds are polyatomic ions present in both members of the pair?
    a. SO_3 and $CaSO_4$
    b. NH_4Br and KClO
    c. Cu_2CO_3 and $CuCO_3$
    d. Na_2SO_4 and Na_2S
7. Name each compound in the following pairs of polyatomic ion containing compounds.
    a. $CuNO_3$ and $Cu(NO_3)_2$
    b. $Pb_3(PO_4)_2$ and $Pb_3(PO_4)_4$
    c. $Mn(CN)_3$ and $Mn(CN)_2$
    d. $Co(ClO_3)_2$ and $Co(ClO_3)_3$

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8. Write chemical formulas for the following compounds containing polyatomic ions.
    a. copper(II) sulfate
    b. manganese(III) hydroxide
    c. ammonium nitrate
    d. magnesium phosphate
9. Name the following binary molecular compounds.
    a. S_4N_2
    b. SO_3
    c. IF_7
    d. N_2O_4
10. Write chemical formulas for the following binary molecular compounds.
    a. disulfur monoxide
    b. tetraphosphorus hexoxide
    c. carbon dioxide
    d. silicon tetrachloride
11. Name each of the following compounds as acids.
    a. $HClO_4$
    b. $HClO_3$
    c. $HClO_2$
    d. HClO
    e. HCl
12. Calculate the molar mass (to 0.01 g/mol) of each of the following substances.
    a. $NaHCO_3$ (baking soda, or sodium bicarbonate)
    b. Tl_2SO_4 (thallium(I) sulfate, a rat and ant poison)
13. Calculate the percent element composition for each of the following compounds. (Round off all atomic masses to 0.01 g/mol before using them.)
    a. $C_{10}H_8$ (naphthalene, ingredient in some mothballs)
    b. NaCN (sodium cyanide, used to extract gold from ores)
14. A sample of the compound sodium azide, NaN_3, contains 57.08 g of Na. How many grams of N does it contain?
15. How many particles (atoms, molecules, formula units or ions) are present in 1.00 mole of each of the following?
    a. copper (Cu) atoms
    b. ammonia (NH_3) molecules
    c. phosphate (PO_4^{3-}) ions
16. Calculate the number of molecules present in each of the following samples of molecular compounds.
    a. 4.69 moles CO
    b. 0.433 mole SO_3
17. Calculate the mass, in grams, of 0.981 mole of each of the following compounds.
    a. SO_2
    b. S_4N_4

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18. A 0.571 mole sample of a pure substance has a mass of 36.60 g. What is the molal mass of the substance?
19. Calculate the number of atoms present in a 3.752 g sample of each of the following elements.
    a. Li
    b. V
20. Each of the following is a correctly written molecular formula. In each case write the empirical formula for the substance.
    a. As_4O_6
    b. Pb_3S_4
    c. C_4H_8
21. Given the following percent element compositions, determine the empirical formula.
    a. 47.26% Cu and 52.74% Cl
    b. 40.27% K, 26.78% Cr, and 32.96% O
22. Determine the molecular formulas of compounds with the following empirical formulas and molecular masses.
    a. CB_2H_3, 73.3 amu
    b. $C_5H_{10}O_2$, 102 amu
    c. $SNCl_2$, 351 amu
23. Adipic acid, a compound used in the manufacture of nylon, has a molecular mass of 146 g/mol. Its percent composition by mass is 49.30% C, 6.9l% H, and 43.79% O. Determine the molecular formula of adipic acid.

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